

Friday, October 4/13  
Science 10

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1. Periods/Groups
  2. Period/Group Numbers
  3. Group Characteristics
  4. **Test: Chapters 3 and 4 -> Tuesday, Oct. 8/13**
  5. Test Topics - See the next page in this attachment. Pick up a hard copy at lunch if you can't access the school website from home.
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6. Activity - Exploring the Modern Periodic Table (Page 108)



## Test - Text: Science 9 -> C3 and C4

### Chapter 3

- Rutherford's Nuclear Model
  - nucleus surrounded by empty space
- subatomic particles: protons (+)  
neutrons (neutral)  
electrons (-)
- atomic number = number of protons 

#
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- mass number = #p + #n
- standard atomic notation  ${}_{17}^{35}\text{Cl}$
- Bohr's Planetary Model of an Atom - 5 statements
- Bohr Diagrams and Bohr-Rutherford Diagrams
- ions (net charge, name changes)
- isotopes, radioisotopes and three types of radioactive decay

### Chapter 4

- periodic table
- Mendeleev -> arranged elements to create the Periodic Table
- Mendeleev's and Modern Periodic Laws
- be able to label the group and period names, period and group numbers, color the metals/metalloids/nonmetals
- group # = # of valence electrons
- period # = # orbits
- characteristics of alkali metals, halogens, noble gases and hydrogen
- trends: boiling points, density, atomic radius, etc.

## Attachments

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