

Monday, October 21/13
Science 10

1. Blackline Master 5.8 - Ionic Compounds - **Pass in for Marking**
 2. Polyatomic Ions
 3. Polyatomic Compounds
 4. **Understanding Concepts - Page 198, #1-4, 6-7 -> HW**
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5. Oxyacids
6. Ionic Compounds vs Molecular Compounds



Page 198 - Do #1-4, 6-7.

Understanding Concepts

- In your own words, explain what is meant by the term "polyatomic ion." Give two examples.
- What happens to the ions in the compound sodium nitrate when it dissolves in water?
- Write the formulas for the following compounds:
 - sodium phosphate
 - calcium sulfate
 - potassium chlorate
 - aluminum hydroxide
 - beryllium nitrate
 - magnesium hydrogen carbonate (magnesium bicarbonate)
 - nickel carbonate
- Write the names for the following compounds:
 - K_2CO_3
 - Na_2SO_4
 - $Al(HCO_3)_3$
 - $AgNO_3$
- What pattern do you see in the formulas of the oxyacids and the original ionic charge of the polyatomic ion? Explain, with two examples.
 - Why does this pattern make sense?
- Why is ammonium nitrate (NH_4NO_3) not written as $N_2H_4O_3$?
- Some polyatomic ions have a positive charge. The ammonium ion (NH_4^+) is an example. Give the names and formulas of the compounds formed by this ion with:
 - a chloride ion
 - a sulfate ion