

Wednesday, October 16/13
Science 10

1. IS -> Intervention: Thursday and Friday (This Week)
 - > My Classroom
 - > Test: C3 and C4 (Text: Science 9)
2. Naming Ionic Compounds - Continue
3. Multivalent Metals
4. Understanding Concepts - Page 195, #1-10
5. Polyatomic Ions
6. Polyatomic Compounds
7. Understanding Concepts - Page 198, #1-4, 6-7



Page 195 - Do #1-10

Understanding Concepts

- (a) How does the sum of the charges on the positive ions compare to the sum of the charges on the negative ions in ionic compounds?

(b) Calculate the sum of the ionic charges in the compound Al_2O_3 . Show your calculation.
 - Draw a Bohr diagram to show the electron transfer that occurs when magnesium and fluorine form the compound magnesium fluoride.
 - Write the formulas for the compounds formed by the following combinations of elements:
 - lithium and fluorine
 - calcium and bromine
 - sodium and nitrogen
 - aluminum and nitrogen
 - Name each of the compounds in question 3.
 - Write the formulas for the following compounds:
 - sodium iodide
 - beryllium fluoride
 - magnesium oxide
 - aluminum sulfide
 - Write the names for the following compounds:
 - KCl
 - Na_3P
 - CaF_2
 - Write the formulas for the following compounds:
 - copper(I) bromide
 - copper(II) bromide
 - iron(II) sulfide
 - Write the names for the following compounds:
 - SnCl_2
 - SnCl_4
 - PbBr_2
 - Write the formula and name of the compound formed by each of the following combinations of ions. (Note that some of these ions will require the use of Roman numerals in the names.)
 - Fe^{3+} and O^{2-}
 - Ca^{2+} and F^-
 - Cu^+ and S^{2-}
 - In mining, some minerals are referred to as ferrous. What metallic element is present in these compounds? (Hint: Look at the letters that begin the word.)
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