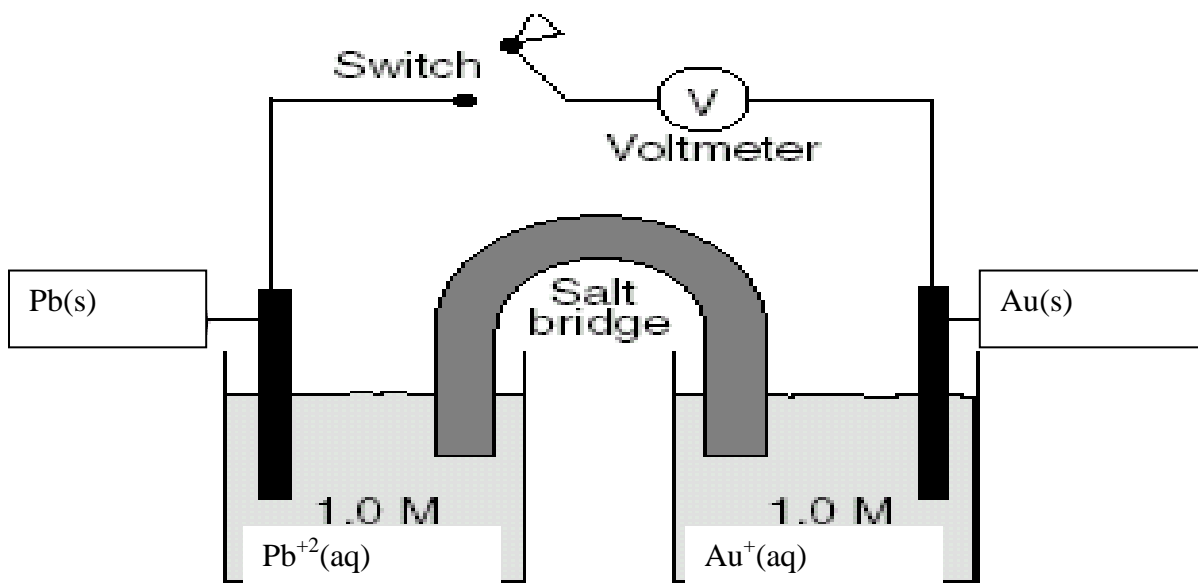


### Voltaic Cells, Example #3

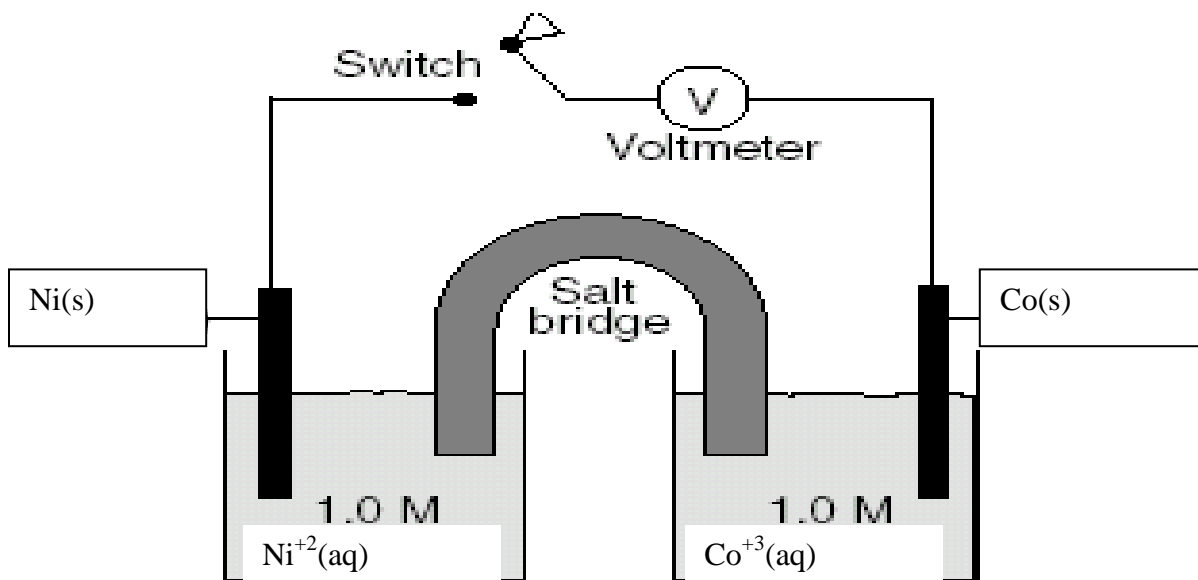


Half Reactions: \_\_\_\_\_

Balanced Net Reaction: \_\_\_\_\_

1. Identify the direction of electron flow \_\_\_\_\_
2. Identify the direction in which + ions will move \_\_\_\_\_
3. Circle the anode on the diagram.
4. Put a box around the cathode on the diagram.
5. Label each electrode whether it is + or -
6. Which electrode will decrease in mass? \_\_\_\_\_
7. Which ion concentration will increase? \_\_\_\_\_

### Voltaic Cells, Example #4



Half Reactions: \_\_\_\_\_

Balanced Net Reaction: \_\_\_\_\_

- 1 Identify the direction of electron flow \_\_\_\_\_
2. Identify the direction in which + ions will move \_\_\_\_\_
3. Circle the anode on the diagram.
4. Put a box around the cathode on the diagram.
5. Label each electrode whether it is + or -
6. Which electrode will decrease in mass? \_\_\_\_\_
7. Which ion concentration will increase? \_\_\_\_\_