

# Physics 112

Monday, June 6/16

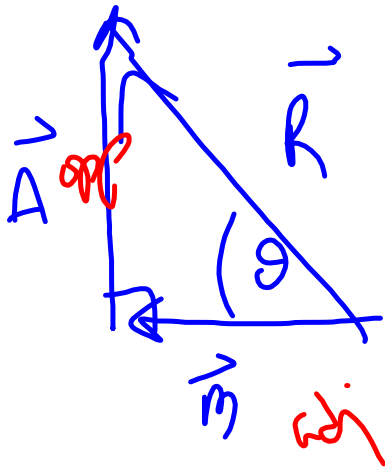
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1. Return Marks -> Test - Unit 3
  2. Reports
  3. Exam Review
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## Physics 112 - Exam Review

Find  $\vec{R}$  analytically.Find the resultant of  $\vec{A} = 67 \text{ N, N}$  and  $\vec{B} = 51 \text{ N, W}$ .Sketch.① Magnitude

$$R^2 = A^2 + B^2$$

$$R = \sqrt{(67)^2 + (51)^2}$$

$$R = \underline{84 \text{ N}}$$

② Direction~~84 N, 53° N of W~~

$$\vec{R} = 84 \text{ N, } 53^\circ \text{ N of W}$$

$$\vec{R} = 84 \text{ N } [53^\circ \text{ N}] \text{ ret.}$$

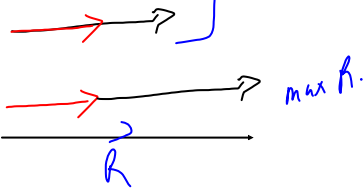
$$\tan \theta = \frac{A}{B}$$

$$\tan \theta = \frac{67}{51} \quad \text{2SD.}$$

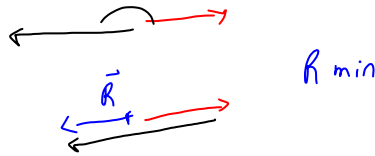
$$\theta = \underline{53^\circ} \quad \text{2SD.}$$

MC  $\rightarrow$  Min/Max R values.

$0^\circ$  - between



$180^\circ$  - between



$\rightarrow$  which angle produces the largest R.

- a)  $0^\circ$
  - b)  $40^\circ$
  - c)  $90^\circ$
  - d)  $180^\circ$
- b) smallest R*

Conversions.

cm  $\leftrightarrow$  m  
 1m = 100 cm  $\times$   
 1cm =  $10^{-2}$  m

$\xleftarrow{-3.6}$   
 m  $\leftrightarrow$  Km  
 $\xrightarrow{\times 3.6}$   
 s  $\leftrightarrow$  h

Km  $\leftrightarrow$  m

KW  $\leftrightarrow$  W

1000m = 1Km  
 1000W = 1KW

1min = 60s

1h = 3600s

ms  $\leftrightarrow$  s  
 $\xleftarrow{-3}$   
 1ms =  $10^{-3}$  s

ns  $\leftrightarrow$  s  
 $\xleftarrow{-9}$   
 1ns =  $10^{-9}$  s

g  $\leftrightarrow$  Kg  
 1000g = 1Kg

## Physics 112 - Exam Review

Kinematic Equations

how objects move.

$$\vec{v} = \frac{d\vec{r}}{dt}$$

uniform motion } constant velocity

$$\vec{v}_f = \vec{v}_i + \vec{a}t$$

$$\vec{d} = \frac{1}{2}(\vec{v}_i + \vec{v}_f)t$$

$$\vec{d} = \vec{v}_i t + \frac{1}{2}\vec{a}t^2$$

$$v_f^2 = v_i^2 + 2\vec{a}d$$


uniformly acc'd motion  
 → acc. is constant  
 → vel. is changing.

Freely Falling Body Problem

June 2016

## Science 122

Monday, June 6/16

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1. Return Test - Nuclear Physics

2. Progress Reports

3. Worksheet #63

Worksheet #64

Worksheet - Assigning Oxidation Numbers

Worksheet - Redox Reactions and Electrochemistry

4. Assignment - Electrochemistry -> Wed., June 8/16.

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## Science 10

Monday, June 6/16

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1. Article - Risks Feared in Bioengineered Insects

- 1 Day Late

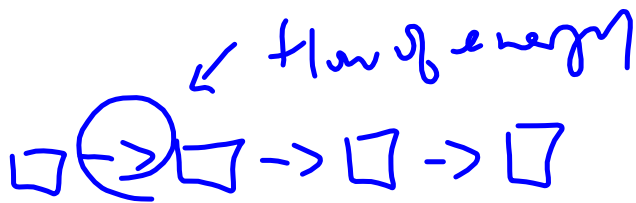
2. Assigment - Ecology -> Wed.

3. Exam Review

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## Science 10 - Last Assignment!

- ecology *- def*
  - ecological levels in an ecosystem:
    - organism
    - species
    - population
    - community
    - ecosystems
    - ecotones
    - biomes
    - biosphere
- Simple.*
- ↓
- Complex*
- indicator species
  - abiotic and biotic factors
  - organisms:
    - producer
    - consumer
      - herbivore
      - carnivore
      - omnivore
    - decomposer
  - food chain
  - detritus food chain
  - primary energy source (sun)
  - photosynthesis (word and balanced chemical equations)
  - trophic levels:
    - primary producer (autotroph)
    - primary consumer (1st heterotroph - herbivore)
    - secondary consumer (2nd heterotroph - 1st carnivore)
    - tertiary consumer (3rd heterotroph - 2nd carnivore)
    - quaternary consumer (4th heterotroph - top carnivore)
  - food web

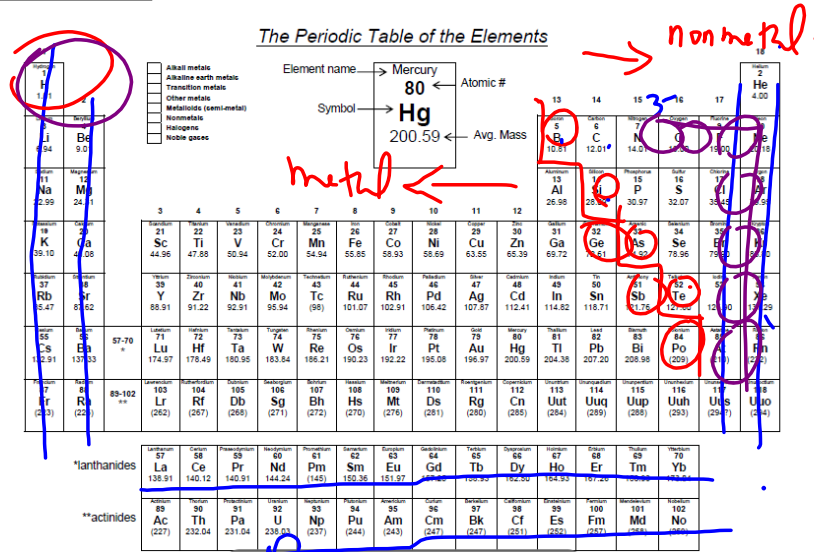


- biodiversity
- bioaccumulation and biomagnification (ie/ mercury, DDT)
  - toxin
  - dosage
  - duration
- sustainability
- types of substances: organic (proteins, sugars, fats)  
inorganic ( $\text{CO}_2$ ,  $\text{H}_2\text{O}$ ,  $\text{NH}_3$ )
- movement of energy ( $\longrightarrow$ ) and matter ( $\bigcirc$ ) through ecosystems
- matter and atoms flows cycles
- biogeochemical cycles
- reservoirs - atmosphere, oceans, earth  
 $\text{N}_2 > \text{O}_2 > \text{C}$  atm. sphere.
- carbon cycle - key element in living things
  - photosynthesis and cellular respiration ✓  
(word and balanced chemical equations)
  - concept map
- nitrogen cycle - needed to make proteins and DNA
  - nitrogen fixation ( $\text{N}_2 \rightarrow \text{NO}_3^-$ )
  - denitrification ( $\text{NO}_3^- \rightarrow \text{N}_2$ )



# Science 10 - Exam Review

## Periodic Table



### Element/Atoms/Ions

Element Name	Element Symbol	Atomic Number	# e <sup>-</sup> Atom	Ion Name	# e <sup>-</sup> Ion	Ion Symbol
sodium	Na	11	11	Sodium ion	10	Na <sup>+</sup>
fluorine	F	9	9	fluoride ion	10	F <sup>-</sup>
neon	Ne	10	10			

Handwritten notes: "neutral", "#p+", "metal", "NH4<sup>+</sup>".

### Ionic Compounds

- transfer of e<sup>-</sup>'s. (metal, NH4<sup>+</sup>)
- electrically neutral
- binary ionic cpds. NaCl (name)
- polyatomic ions NH<sub>4</sub>SO<sub>4</sub> (sulfate, NH4<sup>+</sup>)
- multi-valent metal CuSO<sub>4</sub> (sulfate, Cu)
- copper(II) sulfate

### Molecular Compounds

- sharing of e<sup>-</sup>'s. (NH<sub>3</sub>)
- prefixes. (metalloid)
- CO<sub>2</sub> → carbon dioxide
- diatomic molecules: O<sub>2</sub>, N<sub>2</sub>, F<sub>2</sub>, Br<sub>2</sub>, Cl<sub>2</sub>, I<sub>2</sub>
- S<sub>8</sub> → sulfur P<sub>4</sub> → phosphorus (chlorine...)

### Chemical Reactions

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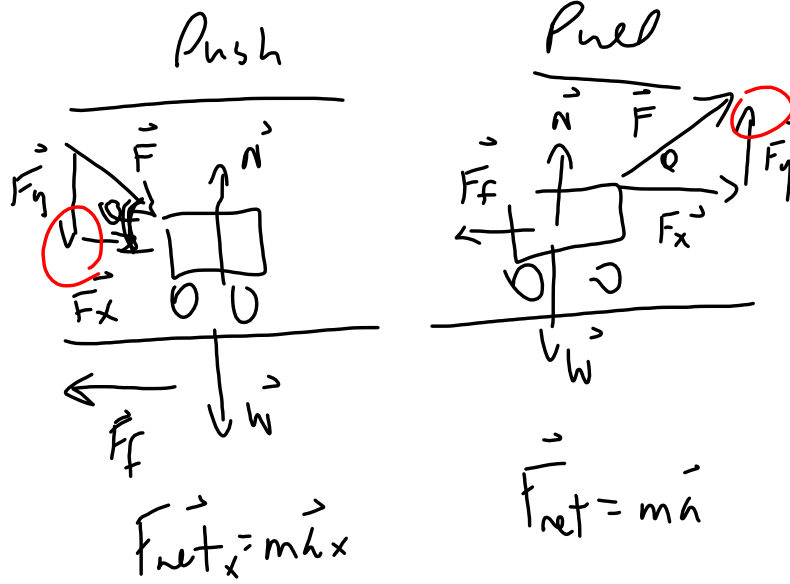
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1. Exam Topics
  2. Marks -> Quiz - Electrostatics
  3. Worksheet - Textbook: Page 719, C15 - PP #27-31 (S. Circuits)  
Textbook: Page 724, C15 - PP#32-35 (P. Circuits)

Textbook: Page 728, PP #36-37 } (C. Circuits)  
Textbook: Page 749, PFU #33-34 }

4. Exam Review
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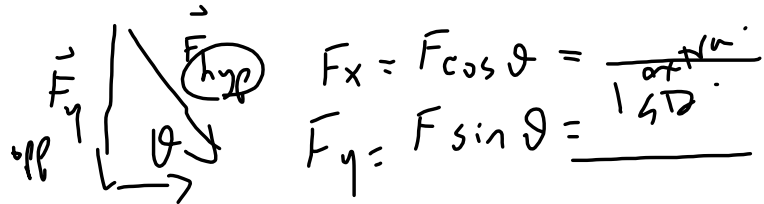
Physics 122 - Exam Review

Push/Pull Problems



$+F_x - F_f = 0$  [const. vel]

$+F_x - F_f = m(a)$  [acc]



$F_f = \mu N$

$F_f = \mu(mg + F \sin \theta)$

$N \neq W$

$F_{net,y} = m\dot{a}_y$

$+N - W - F_y = 0$

$N = W + F_y$

$N = mg + F \sin \theta$

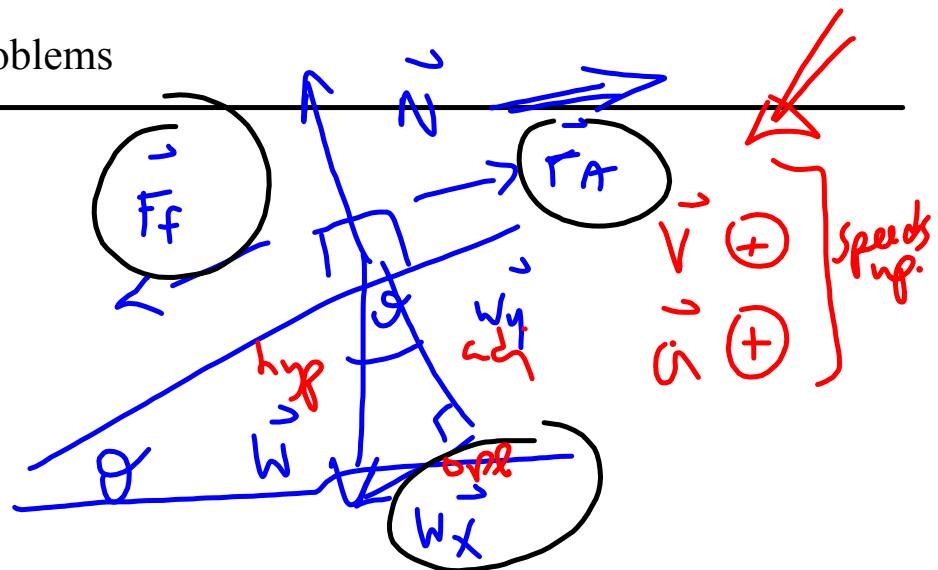
Push

$F \cos \theta - \mu(mg + F \sin \theta) = ma$

$F \cos \theta - \mu mg - \mu F \sin \theta = ma$

## Physics 122 - Exam Review

## Inclined Plane Problems



$$F_{net\ x} = ma_x$$

$$+F_A - F_f - W_x = m(+a)$$

$$F_A - \mu N - W_x = ma$$

$$F_A - \mu W_y - W_x = ma$$

$$F_A - \mu W \cos \theta - W \sin \theta = ma$$

$$F_A - \mu mg \cos \theta - mg \sin \theta = ma$$