<u>Counting Atoms</u>

NaCl	Na	1
	Cl	1
MgF_2	Mg	1
	F	2
K_3PO_4	K	3
	P	1
	O	4
$Sr(OH)_2$	Sr	1
	O	2
	H	2

} Brackets

$Al_2(SO_4)_3$	Al	2
	S	3
	O	12
2RbI	Rb	2
	I	2
$4Ca_3P_2$	Ca	12
	P	8

} Coefficient

|		
$3Zn(NO_3)_2$	Zn	$3 \times 1 = 3$
	N	$3 \times 1 \times 2 = 6$
	O	$3 \times 3 \times 2 = 18$

★

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Worksheet: Counting Atoms in Compounds

Oct 2-12:00 AM

Chapter 6 - Understanding Chemical Reactions

(Page 216)

A **chemical reaction** is a process that leads to the transformation of one set of **chemical** substances to another.

substance + substance $\xrightarrow{\text{Yeild}}$ substance + substance

reactants
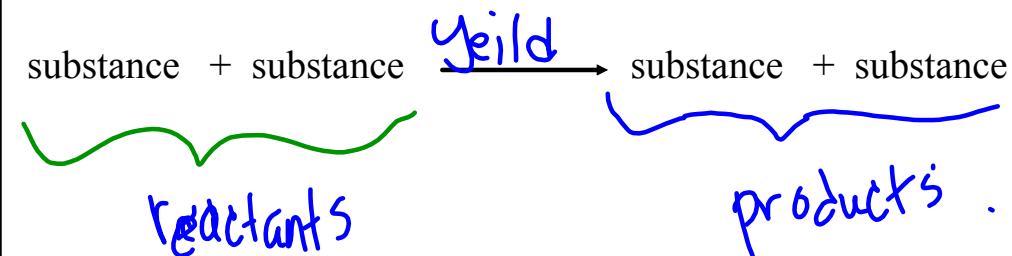
products .

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Conserving Mass

The Law of Conservation of Mass states that, in a chemical reaction, the total mass of the reactants is always equal to the total mass of the products.

mass reactants = mass products

Mar 30-11:40 PM

<u>Counting Atoms</u>

NaCl	Na
	Cl
MgF$_2$	Mg
	F
K$_3$PO$_4$	K
	P
	O
Sr(OH)$_2$	Sr
	O
	H
Al$_2$(SO$_4$)$_3$	Al
	S
	O
2RbI	Rb
	I
4Ca$_3$P$_2$	Ca
	P
3Zn(NO$_3$)$_2$	Zn
	N
	O

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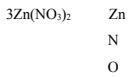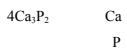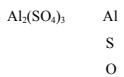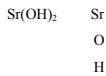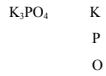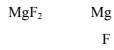
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<u>Balancing Chemical Reactions</u>

1. ___ N$_2$ + _3_ H$_2$ ⟶ _2_ NH$_3$

   N 2
   H 2 6

   N 1 2
   H 3 6

2. ___ P$_4$ + _3_ O$_2$ ⟶ _2_ P$_2$O$_3$

3. _2_ K + ___ MgBr$_2$ ⟶ _2_ KBr + ___ Mg

4. __ FeCl$_3$ + _3_ NaOH ⟶ __ Fe(OH)$_3$ + _3_ NaCl
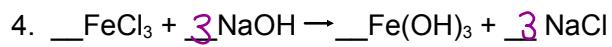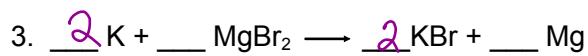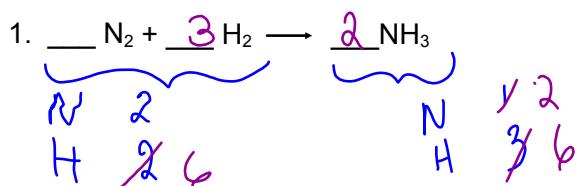
Apr 7-12:15 PM

Worksheet - Balancing Chemical Reactions

http://misterguch.brinkster.net/WKS001_019_348432.pdf

Oct 2-12:25 AM

Attachments

Science 10 - Grade 9 Chem Topics.docx

Science 10 - Grade 9 Chem - What Do You Know.docx

Science 10 - Activity - Molecular Models.docx

Science 10 - Answer Key - Ions and Subatomic Particles.pdf