

Science 122

Monday, October 3/16

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1. Summative Assessment - Magnetism

Physics 112

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1. Questions?
Summative Assessment - U1: S1 and S2
- **Wednesday, Oct. 5/16**
2. Midterm - Wednesday, Nov. 9/16
3. Uniformly Accelerated Motion: Kinematic Equation #1
4. Uniformly Accelerated Motion: Kinematic Equation #2
5. Uniformly Accelerated Motion: Kinematic Equation #3
6. Uniformly Accelerated Motion: Kinematic Equation #4
- To Be Continued

7. Quadratic Formula
8. Worksheet - Motion Problems

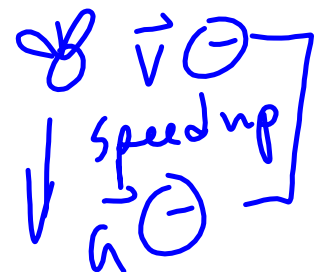
Topics: Unit 1 - S1 and S2

Section 1 - Vector Analysis

1. mechanics, kinematics, dynamics *d efⁿ*
2. types of physical quantities:
 - (i) scalar quantity - has magnitude only
 - examples *d, v, t, m*
 - (ii) vector quantity - has magnitude and direction
 - examples *\vec{v} , \vec{a} , $\Delta\vec{d}$, \vec{J}*
 - conventional directions
 - vector notation
 - graphical representation \rightarrow
3. resultant = vector sum
4. graphical addition of vectors:
 - (i) tip-to-tail method *(Lead-to-tail)* 7 km
 5 km
 - (ii) parallelogram method
5. range of resultant magnitudes *minR — maxR*
6. calculate a resultant (follow rubric)
7. types of motion:
 - (i) no motion 180° 0°
 - (ii) uniform motion *\vec{v} const.* 2 km — 12 km
 - (iii) uniformly accelerated motion *\vec{a} const. / \vec{v} changing "*
8. use directions of velocity and acceleration to describe motion

Section 2 - Graphical Analysis

1. position-time graphs -> interpret
2. position-time graph -> direction of motion
3. velocity-time graph -> interpret
4. velocity-time graph -> direction of motion
5. velocity-time graph -> calculations



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1. Experiment 5.2 - Friction
 - Due: Friday, Sept. 30/16
 - **1 Day Late**

2. Check -> Worksheet - Static Torque #1

3. Static Torque - Forces Acting at Angles

4. Worksheet - Static Torque #2

5. Experiment 10.2 - Torques (Page 67)

Science 10

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1. September Progress Reports - Take Home
 - Have a Parent/Guardian Sign It
 - **Return by Tuesday, Oct. 4/16**
 2. Questions?
 - Test #1 - Chemistry to the End of Compounds - Tuesday, Oct. 4/16**
 3. Balancing Chemical Equations - Continue
 4. Worksheet - Balancing Chemical Equations - Not HW
-
5. Types of Chemical Reactions
 6. Formation/Synthesis Reactions - General Format
 - Examples
 7. Decomposition Reactions - General Format
 - Examples
 8. Worksheet - Formation and Decomposition Reactions
 9. Single Replacement Reactions
 10. Double Replacement Reactions

Topics

Test #1 - Chemistry to the End of Compounds

1. chemistry
2. periodic table of the elements - rows -> periods
 - columns -> groups/families
 - family and period names
 - chemical symbols
 - location of metals, nonmetals and metalloids
3. atoms -> building blocks of matter
 - > three subatomic particles: p^+ , n , e^-
 - > locations of three subatomic particles
 - > electrically neutral: $\#p^+ = \#e^-$
4. atomic number = number of protons
5. characteristics of metals and nonmetals
6. ions - atoms that have gained or lost electrons
 - cations/positive ions/metallic ions
 - anions/negative ions/nonmetallic ions
 - be able to state number of protons, number of electrons and ion charges
7. ionic bond - created by transfer of electrons
8. be able to identify monatomic ions, polyatomic ions and monatomic ions of multivalent metals
9. ionic compounds - electrically neutral
10. be able to write the names of simple binary ionic compounds given their formulas and vice versa
11. be able to write the names of ionic compounds containing polyatomic ions given their formulas and vice versa
12. roman numerals 1-10
13. be able to write the names of ionic compounds containing multivalent metals given their formulas and vice versa
14. be able to write the names of ionic compounds containing multivalent metals and polyatomic ions given their formulas and vice versa
15. covalent bond - created as a result of the sharing of electron pairs
16. molecular compounds = covalent compounds = molecules
17. prefixes 1-10
18. diatomic molecules: H_2 , N_2 , O_2 , F_2 , Cl_2 , Br_2 , I_2
19. special molecules: P_4 , S_8 , water, ammonia, hydrogen peroxide
20. be able to write the names of binary molecular compounds given their formulas and vice versa
21. identify ionic compounds and molecular compounds