# Science 10

Tuesday, February 28/17

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1. Assignment - All Ionic Compounds
   - Wednesday, March 1/17
   - Topics
   - Sample Questions

2. Worksheet #5 - Ionic Compounds Summary
   Worksheet - Lots of Ionic Naming Practice Problems

3. Covalent Bonds

4. Diatomic Molecules

5. Naming Binary Molecular Compounds

6. Worksheet - Practice: Binary Covalent Compounds

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7. Recap Types of Compounds

8. Worksheets - Mixed Ionic and Covalent Compounds

# Science 10
## Topics -> Assignment: All Ions and Ionic Compounds

1. a) be able to identify monatomic ions

   b) be able to write the names of monatomic ions given their chemical symbols and vice versa

   $Cl^-$ → chloride ion  Sodium ion → $Na^+$

2. be able to write the names of simple binary ionic compounds given their formulas and vice versa

   NaCl ⇄ Sodium chloride

3. a) be able to identify polyatomic ions by their symbols and names ("ate", "ite" and some "ide" endings)

   sulfate ion  nitrite ion  hydroxide ion

   b) know where to find the names and symbols of polyatomic ions on the purple periodic table

   c) be able to write the names of ionic compounds containing polyatomic ions given their formulas and vice versa

   $NaNO_3$ ⇄ Sodium nitrate

4. a) be able to identify multi-valent metals

   copper (I)  $Cu^{+2}$ / $Cu^+$

   b) be able to write the names of multivalent metal ions using roman numerals given their symbols and vice versa

   1 to 10  I ⟶ X

   c) be able to write the names of ionic compounds containing multivalent metals given their formulas and vice versa

   Worksheet #4

5. be able to write the names of ionic compounds containing multivalent metals and polyatomic ions given their formulas and vice versa

   Copper (I) sulfate ⇄ $Cu_2SO_4$

## Assignment - All Ionic Compounds -> Sample Questions

### Types of Ions

Identify each of the following as a monatomic ion (MI), a polyatomic ion (PI), or the ion of a multivalent metal (IMM), by printing MI, PI or IMM on the line provided.

a) $F^-$ MI

b) lead (II) ion IMM

c) $CNO^-$ PI

d) sulfite ion PI

e) $Pt^{2+}$ IMM

| MI | PI | IMM |
|---|---|---|
| $NH_4^+$ | -ate | roman numeral |
| $Cl^-$ | -ite | I, II, V |
| | -ide | |
| | 3 | |

$NO_3^-$ $Cu^{+2}$ $Cu^+$

### Chemical Formulas to Chemical Names

Name the following ionic compounds.

a) $K_2S$ potassium sulfide

b) $SnBr_4$ tin (IV) bromide

c) $Zn(ClO_4)_2$ zinc perchlorate $ClO_4^-$

### Chemical Names to Chemical Formulas

Write the chemical formulas for the following ionic compounds.

a) barium cyanide $Ba^{2+}(CN)^- \Rightarrow Ba(CN)_2$

b) ammonium phosphide $(NH_4)_3P$

c) gold (III) chloride $Au^{3+}Cl^-_3$ $AuCl_3$

d) magnesium borate $Mg^{2+}_3(BO_3^{3-})_2$ $Mg_3(BO_3)_2$

Worksheet #5
Lots of Practice

# Physics 112

Tuesday, February 28/17

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1. Return - SA - U1 S1
2. Position-Time Graphs - To Be Continued
3. Position-Time Graph: Direction of Motion
4. Velocity-Time Graphs
5. Velocity-Time Graph: Direction of Motion
6. Velocity-Time Graph Calculations - To Be Continued

7. Worksheet: Velocity-Time Graph #1-4

# Physics 122

Tuesday, February 28/17

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1. SA - U1 S1 - 3 Problems (40 minutes)
   - Wednesday, March 1/17
2. Unit 1 - Section 2 -> Torque
3. Center of Mass
4. Types of Motion - Translational and Rotational
5. Torque - To Be Continued

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6. Net Torque
7. Static Equilibrium - Revisited