

## Science 10

Tuesday, February 28/17

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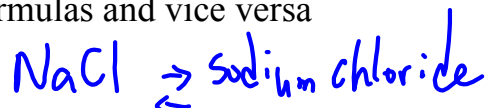


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1. Assignment - All Ionic Compounds
    - **Wednesday, March 1/17**
    - Topics
    - Sample Questions
  2. Worksheet #5 - Ionic Compounds Summary  
Worksheet - Lots of Ionic Naming Practice Problems
  3. Covalent Bonds
  4. Diatomic Molecules
  5. Naming Binary Molecular Compounds
  6. Worksheet - Practice: Binary Covalent Compounds
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7. Recap Types of Compounds
  8. Worksheets - Mixed Ionic and Covalent Compounds

## Science 10

## Topics -&gt; Assignment: All Ions and Ionic Compounds

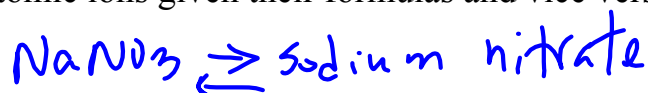
1. a) be able to identify monatomic ions
  - b) be able to write the names of monatomic ions given their chemical symbols and vice versa
2. be able to write the names of simple binary ionic compounds given their formulas and vice versa



3. a) be able to identify polyatomic ions by their symbols and names ("ate", "ite" and some "ide" endings)



- b) know where to find the names and symbols of polyatomic ions on the purple periodic table
- c) be able to write the names of ionic compounds containing polyatomic ions given their formulas and vice versa



4. a) be able to identify multi-valent metals

copper (I)



- b) be able to write the names of multivalent metal ions using roman numerals given their symbols and vice versa

1 to 10



- c) be able to write the names of ionic compounds containing multivalent metals given their formulas and vice versa

Worksheet #4

5. be able to write the names of ionic compounds containing multivalent metals and polyatomic ions given their formulas and vice versa



**Assignment - All Ionic Compounds -> Sample Questions**

**Types of Ions**

Identify each of the following as a monatomic ion (MI), a polyatomic ion (PI), or the ion of a multivalent metal (IMM), by printing MI, PI or IMM on the line provided.

|                     |            |                 |              |                            |
|---------------------|------------|-----------------|--------------|----------------------------|
| a) F <sup>-</sup>   | <u>MI</u>  | MI              | PI           | IMM                        |
| b) lead (II) ion    | <u>IMM</u> |                 | -ate<br>-ite | Roman numeral.<br>I, II, V |
| c) CNO <sup>-</sup> | <u>PI</u>  | Nh <sup>+</sup> | -ide         |                            |
| d) sulfite ion      | <u>PI</u>  | Cl <sup>-</sup> | 3            |                            |
| e) Pt <sup>2+</sup> | <u>IMM</u> |                 |              |                            |

NO<sub>3</sub><sup>-</sup>      Cu<sup>2+</sup>  
Cu<sup>+</sup>

**Chemical Formulas to Chemical Names**

Name the following ionic compounds.

a) K<sub>2</sub>S      potassium sulfide

b) SnBr<sub>4</sub>      tin (IV) bromide

c) Zn(ClO<sub>4</sub>)<sub>2</sub>      zinc perchlorate      ClO<sub>4</sub><sup>-</sup>

**Chemical Names to Chemical Formulas**

Write the chemical formulas for the following ionic compounds.

a) barium cyanide      Ba<sup>2+</sup> (CN)<sup>-</sup> → Ba(CN)<sub>2</sub>

b) ammonium phosphide      (NH<sub>4</sub>)<sub>3</sub>P

c) gold (III) chloride      Au<sub>1</sub>Cl<sub>3</sub>      AuCl<sub>3</sub>

d) magnesium borate      Mg<sup>2+</sup> (BO<sub>3</sub>)<sup>3-</sup> → Mg<sub>3</sub>(BO<sub>3</sub>)<sub>2</sub>

Worksheet #5  
Lots of practice

Mg<sub>3</sub>(BO<sub>3</sub>)<sub>2</sub>

## Physics 112

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1. Return - SA - U1 S1
  2. Position-Time Graphs - To Be Continued
  3. Position-Time Graph: Direction of Motion
  4. Velocity-Time Graphs
  5. Velocity-Time Graph: Direction of Motion
  6. Velocity-Time Graph Calculations - To Be Continued
- 
7. Worksheet: Velocity-Time Graph #1-4

## Physics 122

Tuesday, February 28/17

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1. SA - U1 S1 - 3 Problems (40 minutes)  
- Wednesday, March 1/17
  2. Unit 1 - Section 2 -> Torque
  3. Center of Mass
  4. Types of Motion - Translational and Rotational
  5. Torque - To Be Continued
- 
6. Net Torque
  7. Static Equilibrium - Revisited