Science 10 **PRACTICE EXAM Chemistry and Physics**

Name -

INSTRUCTIONS

- 1. Write your first and last name on the line provided above.
- 2. All parts of the practice exam will be completed on this paper.

Part 1 - Multiple Choice (Value - 30)

<u>Circle</u> the letter of the best answer.

X or ÷

1. The Certainty Rule is used to

3. Which list includes only metals?

(a) krypton, nitrogen, helium, xenon (b) titanium, zinc, copper, chlorine (c) gold, mercury, boron, iron

(d) nickel, platinum, lead, aluminum

- (a) determine the number of significant digits when adding and subtracting measured values
- (b) determine the number of significant digits when adding and multiplying measured values
- (c) determine the number of significant digits when subtracting and dividing measured values
- (d) determine the number of significant digits when multiplying and dividing measured
- 2. You are riding your bicycle west. If you decide to decrease your velocity (slow down), in what direction is your acceleration?
 - (a) south
 - (b) north
 - (c) east (d) west

4 dir. opposite

no charge

- metals
- 4. Which of the following is electrically neutral?
 - (a) electron
 - (b) proton
 - (c) atom
 - (d) ion

5. A family in the periodic table that contains only nonmetals is the

(a) actinides

- (b) alkali metals
- (c) lanthanides
- (d) halogens

~ -> nonmitals

- 6. Which substance in the following list is an element?
 - (a) NH_3
 - (b) CH₄
 - (c) Bi
 - (d) H₂O
- 7. Which statement is not true about matter? (a) Matter can be destroyed.

(b) Matter is made up of atoms.

- (c) Matter takes up space.
- (d) Mattan ha

9. Molecular compounds are formed when electrons are

(a) transferred

(b) shared

(c) lost

(d) gained

10. The measurement 2.030×10^3 km has

(a) 7 significant digits

(b) 2 certain digits

4 certain digits

 \rightarrow (d) l uncertain digit

11. An atom becomes an ion with a charge of -2 when it:

(a) gains 2 protons

(b) loses 2 neutrons

(c) loses 2 electrons

(d) gains 2 electrons

12. Which of the following is an example of speed?

- (a) 40 km
- (b) 20 km/h[E]
- (c) 1.5 m [right]
- (d)15 km/h

Amognitude / no Lirection

13. Which of the following graphs illustrates an object that is speeding up?



Time

- (a) A
- (b) B
- (d) D

14. Which element is a member of the halogens?

- (a) astatine
- (b) copper
- (c) radium
- (d) potassium

15. Protons are

(a) negatively charged particles found outside the nucleus of an atom.

(b) neutral particles found in the nucleus of an atom.

(c) positively charged particles found outside the nucleus of an atom.

(d) positively charged particles found in the nucleus of an atom.

16. The prefix for 5 is

- (a) tri
- (b) di
- © penta
- (d) hexa

pt = #e

17. A beryllium atom has

(a) 4 protons and 6 electrons.

(b) 4 protons and 2 electrons.

(c) 4 protons and 8 electrons.

A materia and 1 1

19. The slope of this position-time graph represents:

(a) position

- (b) time
- (c) velocity
- (d) displacement

rise > m run s



time

- 20. Identify the products in the reaction below.
 - (a) AB and AD (b) CD and CB
 - (c) AB and CD (d) AD and CB
- AB+CD, -> AD+CB (cuctants products
- 21. An element that forms a diatomic molecule is
 - (a) iodine $\mathcal{J}_{\mathcal{D}}$

(b) sulfur

- (c) phosphorous
- (d) sodium
- 22. John walks to his friend's house 5 blocks east and then walks 5 blocks west. His displacement is
 - (a) 10 blocks
 (b) zero blocks
 (c) 10 blocks [E]
 - (d) 10 blocks [W]
 - (d) 10 blocks [w]
- 23. Physics is the study of
 - (a) matter
 - (b) energy
 - (c) how objects move (d) matter and energy
- 24. If $1 \min = 60$ s, then 42 s equals
 - (a) 12 min (b) 0.70 min (c) 178 min (d) 2520 min

425× Imin = U. Jumin 605

25. The SI base unit of mass is

- (a) kg
- (b) m
- (c) s
- (d) m/s

26. Acceleration can best be defined as

- (a) the speed at which an object is travelling at a particular instant
- (b) the displacement of an object divided by time
- (c) the total distance covered over the total time measured
- (d) the rate of change in velocity
- 27. In any chemical reaction, the total mass of the reactants is always equal to the total mass of the products. This is known as the law of:(a) chemical reactions
 - (a) chemical reactions
 - (b) conservation of mass
 - (c) constant proportion (d) conservation of anomal
 - (d) conservation of energy
- 28. Which chemical formula correctly matches the chemical name provided? (a) Ni₂(SO₄₃ - nickel (III) sulfate $0/3 \pm 0.05$

29. A vector quantity has

- (a) direction
- (b) size
- (c) size and direction
- (d) none of the above
- 30. Which of the following is a molecular compound held together by covalent bonds? (a) LiBr

Size=maynithte

- (b) CaO
- (c) H₂O
- (d) Mg_3P_2

Part 2 - Atoms and Ions (Value - 12)

Complete the chart below.

Element Name	Atomic Number	Ion Name	Ion Symbol	# of Electrons in the Ion
radon	86		\times	\times
technitium	F=7# F= 43	technitium ion	T27+	36
phosphorous	15	phosphide ion	P ³⁻	18

Junic Nucl Moleular H20 A metul × use prefixes

Part 3 - Compounds (Value - 16)

a) CrBr₃

b) BeSO₄

c)

d)

e)

f)

All₃

FI₃

1

State whether each compound is ionic or molecular and give the compound name or when chemical formula as required.

T

T

Ionic or molecular

Name or Formula

namin

Chromium (111) bromide beryllium sulfate aluminum isdice fluorine triiodide 3+52- -> Cu253

64 p3- > 64P

dinitrogan tatmo and

cobalt (III) sulfide

gallium phosphide

AA

I

NDI

Part 4 – Identifying and Balancing Reactions (Value – 10)

Balance the following equations and state the type of each reaction: formation (F), decomposition (D), single replacement (SR), double replacement (DR), combustion (C) and neutralization (\mathbb{N}).

	Type
a) $_CrCl_2 \rightarrow _Cr + _Cl_2$	D
b) λ KBr + $Mg(NO_3)_2 \rightarrow \lambda$ KNO ₃ + $MgBr_2$	DR
$M = H_2SO_4 + \underline{2} LiOH \rightarrow \underline{12}SO_4 + \underline{2} H_2O$	\mathbb{N}
d) $\underline{\checkmark}_{Co+\underline{3}}O_2 \rightarrow \underline{2}_{Co_2O_3}$	F
e) $Zn + Na_2SO_4 \rightarrow ZnSO_4 + ZnSO_4$	SR
f) $2C_4H_{10} + 13 O_2 \rightarrow 84 CO_2 + 105 H_2O$	C
85	

Part – Translating Word Equations to Balanced Chemical Equations (Value - 14) Use the following word equations to write balanced chemical equations.

1. Copper metal and phosphorus combine to form copper (II) phosphide.

Cu2+p3-

6 Cut Py -> 2 Cu3P2

2. Beryllium hydroxide and potassium nitrate react to form beryllium nitrate and potassium hydroxide.

Belothin + 2KNU3 -> BelNO3)2+2KOH

3. Zinc metal and chromium (III) nitrate combine to form zinc nitrate and chromium metal.

32n +2Cr(NO3)3 > 3Zn(NO3), +2Cr



+ use

F. 3+

Predict the products then write a balanced chemical equations.

 $1. \underline{2} \operatorname{Fe} + \underline{3} \operatorname{MgSO_3} \rightarrow \operatorname{Fea}(50_3)_3 + 3 \operatorname{Mg}$

Part 7 - Physical Quantities (Value - 9)

Physical Quantity	Type of Physical Quantity (scalar or vector)	Variable	Unit (s, m, m/s, m/s ²)
constant velocity	Vector	ר ע	mis
distance	scular	2	m
time	Sculat	t	5
acceleration	Vector	2	52
displacement	Vector	ΔJ	η
average speed	Sculu(VaJ	m/5

Complete the chart below. Choices are provided in some of the headers.

Part 8 – Position, Displacement and Velocity (Value - 24)

- a) Using C as your reference point, state the position of each letter below. Include signs to indicate directions: + for a positive direction and for a negative direction. No units are required in this case. (2)
 - (i) $\mathbf{A} \quad \mathbf{d} = \underline{+5}$ (ii) $\mathbf{B} \quad \mathbf{d} = \underline{-2}$ di dt
 - b) If an object moves from **B** to **A**, what is the object's displacement? Show a calculation. A word statement is not required. *Consult your formula sheet if needed. (3)

 $\begin{aligned} S\vec{J} &= \vec{J} - \vec{J} \\ S\vec{J} &= 5 - (-2) \\ S\vec{J} &= \vec{J} \end{aligned}$



2. Answer the questions below using the position versus time graph provided for an object traveling in a straight line. Assume the positive direction is north.



a) What is the position of the object at t = 10 s? (1)

b) What was the object doing between t = 10 s and t = 15 s? (2)

It Stopped.

c) Did the object change directions? If so, when? (1)

d) What was the maximum displacement of the object? (1)

t=55

e) What type of motion did the object have between t = 0 s and t = 5 s? (1)

uniform motion

25 m = - 5m -55 5

B

 $\frac{35m}{55} = 7m}{3}$

OMEON

f) Draw the velocity-time graph for the object on the grid below. (4)



Answer the questions below using the position versus time graph provided for an object traveling in a straight line. Assume the positive direction is east.



a) What is the velocity of the object at t = 15 s? (1)

6 mls, west

b) (i) What type of motion does the object have between t = 0 s and t = 5s? (1)

no motion

(ii) What type of motion does the object have between t = 15 s and t = 20 s? (1)

hipsonly accelerated notion

c) What did the object do at t = 10 s? (1)

Charged direction (E-SW)

d) What is the acceleration of the object between t = 5 s and t = 15 s? Show a calculation and express your answer to 2 SDs. (2)

$$\vec{a} = \frac{12}{105} \frac{m_{15}}{m_{15}} = -1.2 m_{15}^2 cr 1.2 m_{15}^2 W$$

e) What was the displacement of the object between t = 10 s and t = 15 s? Show a calculation and express your answer to 2 SDs. (3)

$$AJ = 1bh = 1(5)(-6) = -15m$$

Part 9 – Word Problems (Value - 36)

Show work for full value. Include a word statement.

1. A hot air balloon flew 12 h at an average speed of 210 km/h. How far did it travel? (6)

Vau = 210 Km/h	Vau = d	
d=?	J= Vaut	
t = 125	d = (210 km)(12 h)	
		It triveled
	d = 2520 km	2.5×103km.
	$d = 2.5 \times 10^3 \mathrm{km}$	

A race car accelerates at 5.0 m/s² [W]. What was the initial velocity of the race car if its velocity after 2.7 s was 37 m/s [W]? (7)

$$\begin{split} \vec{k} &= -5 \cdot com | s \rangle & \vec{h} &= v \vec{j} - v \vec{i} \quad \forall \vec{j} &= v \vec{i} + \vec{k} \\ \vec{v} &= ? & \vec{k} \\ \vec{v} &= -37 m | s \\ \vec{v} &= -37 m | s \\ \vec{v} &= \vec{v} &= \vec{v} \\ \vec{v} &= \vec{v} &= \vec{i} \\ \vec{v} &= \vec{v} \\ \vec{v} &= -37 - (-5 \cdot c)(2 \cdot 7) \\ \vec{v} &= -34 - (-5 \cdot c)(2 \cdot 7) \\ \vec{v} &= -34 - m | s \\ . \end{split}$$

$$\begin{aligned} \mathcal{R}_{e} \quad initial \quad velocity \quad us \quad 24m | s, \ west \quad . \end{aligned}$$

3. A truck is travelling at 22 m/s east when the driver notices a speed limit sign for the town ahead. If it takes the driver 6.9 s to slow down to a velocity of 14 m/s east, what is the acceleration of the truck? (5)

ら= ジュージョ Vi= + 22mls t= 6.95 $\vec{a} = 14 - (22)$ $\vec{b} \cdot \vec{q}$ if = +14mls エ=フ ã= -1.2 m/52

The haceleration was 1.2 m/s - west

4. A student travels at a constant velocity of 0.38 m/s [E]. How long did it take the student to travel 3.0 m [E]? (6)



- 5. The peregrine falcon is the fastest of the flying birds. If a peregrine falcon can fly 1.73 km downward in 25 s, what is the average velocity of the bird in km/h? (6) $l_h = 36005$
- t = 255 = 0.0069h $V_{AV} = \Delta Jr$ $T_{AV} = -1.73km$ $V_{AV} = -1.73km$ $V_{AV} = -7$ $V_{AV} = -7.5 \times 10^{2} km$ $J_{AV} = -2.5 \times 10^{2} km$

6. An object falls from the Transco Tower in Houston and takes 15 seconds to reach the ground. What is its velocity when it hits the ground? (6)

L= -9. Jum 15 -O t= 155 Vi = omls Ve=7

Vg= V, + at $v_{f} = (-9.80)(10)$ $\vec{v}_{f} = -1.5 \times 10^{2} \text{ m/s}$

Ats velocity was 1.5×10 mls, down