

18.3

SOLUBILITY EQUILIBRIUM

Section Review

Objectives

- Describe the relationship between the solubility product constant and the solubility of a compound
- Predict whether precipitation will occur when the two salt solutions are mixed

Vocabulary

- solubility product constant (K_{sp})
- common ion
- common ion effect

Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

The 1 is the equilibrium constant for the equilibrium between an ionic solid and its ions in solution. The term 2 refers to the lowering of the solubility of a substance by the 3 of a common ion. If the ion-product concentration of two ions in solution is greater than the K_{sp} of the compound formed from the two ions, a(n) 4 will form.

1. Solubility product constant
2. Common ion effect
3. addition
4. precipitate