PROFESSIONAL LEARNING EARLY DISMISSAL DAYS
BELL SCHEDULE

8:35 – 9:07	Period 1
9:07 - 9:12	
9:12 - 9:44	Period 2
9:44 - 9:49	
9:49 – 10:21	Period 3
10:21 -10:26	
10:26-10:58	Period 4
10:58-11:03	
11:03-11:35	Period 5

Physics 112

Wednesday, September 20/17

http://mvhs.nbed.nb.ca/

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1. Learning Targets - Check
2. Review Results of Yesterday's Small Group Activity
3. Rubric: Vector Analysis
4. Examples of Vector Analysis
5. Worksheet - Vector Analysis - Given but not HW

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6. Concepts U1-S2: Graphical Analysis
7. Types of Motion
8. Directions of Velocity and Acceleration
9. Position-Time Graphs

Physics 122

Wednesday, September 20/17

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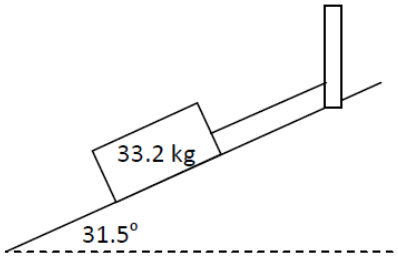
1. Return -> Rated FAs
2. FA - Inclined Plane Problem
3. Worksheets - Force Problems - Type I, II and III (2)

4. FA - Force Problems: Types I, II and III

P122 - Formative Assessment - Incline Problem

The block in the diagram is at rest.
If the coefficient of stac fricon is
0.214, determine the tension in the rope.



Science 10

Wednesday, September 20/17

http://mvhs.nbed.nb.ca/

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1. Assignment - Your Name in Chemical Symbols
   - Due: Friday, Sept. 15/17
   - 3 Days Late

2. FA - Return P5

3. Check -> Worksheet #5 - Ionic Compound Summary

4. Recap: Types of Ions

P5

5. Covalent Bonds

6. Diatomic Molecules

P4

7. Naming Binary Molecular Compounds

8. Worksheets for Duotangs

9. Practice: Binary Covalent Compounds

Science 10

FA – Simple Binary Ionic Compounds and Compounds with Polyatomic Ions (A)

Name - _____ mono - 1 atom.

1. Circle the monatomic ions:

ClO_3^{1-} Mg^{2+} nitrite ion phosphate ion cyanide ion

NO_2^{-1} PO_4^{3-} CN^-

poly poly poly

2. Circle the number of elements involved in an ionic compound involving a polyatomic ion:

1 2 more than 2

$NaNO_3$

3. Write the name for each of the following compound.

a) CaS - _____ Calcium Sulfide

b) Rb_3N - _____ rubidium nitride

c) $Sc(BrO_3)_3$ - _____ Scandium bromate.

polyatomic → green PT.

4. Provide the chemical formula for each compound.

a) ammonium fluoride - _____ $(NH_4)^+ F^- \Rightarrow NH_4F$

poly.

b) sodium phosphide - _____ Na_3P

c) aluminum sulfate - _____ $Al_2(SO_4)_3$

$Al_2^{3+}(SO_4)_3^{2-}$