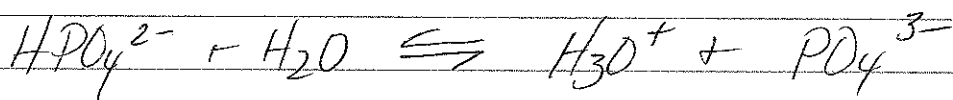
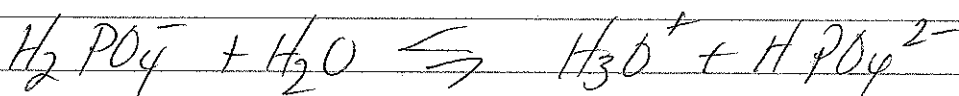
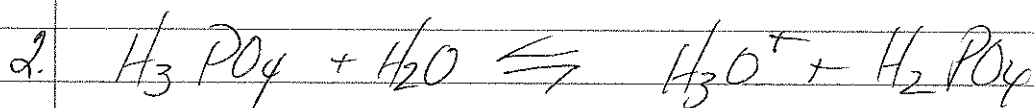
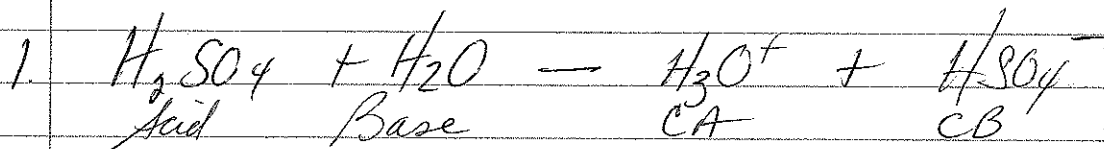


# Practice Problem

19.1

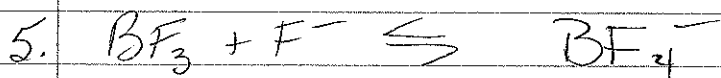
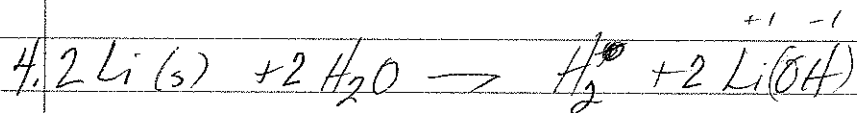


3. a)  $\text{HCOOH}$  - monoprotic

b)  $\text{HBr}$  - monoprotic

c)  $\text{H}_2\text{SO}_4$  - diprotic

d)  $\text{H}_3\text{ClO}_4$  - triprotic



$\text{BF}_3$  - accepts a pair of electrons to form a covalent bond  $\rightarrow$  Lewis Acid

$\text{F}^-$  donates the pair of electrons - Lewis base

6. Sour, tart taste, cause indicators to turn color

Acids react with compounds containing  $\text{OH}^-$  to produce salt +  $\text{H}_2\text{O}$

7. Aqueous sol'n of bases taste bitter or feel slippery, react with acids to produce a salt + water. Base cause indicator to change color