

Science 9

Friday, October 25/19

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School Pictures - Order Deadline: Thursd., Oct. 31

1. The following will be returned next week so those who still need to write can complete them.
 - > Summative Assessment: Calculating Speed, Distance and Time
 - > In-Class Summative Assessment -> The Universe


2. Handouts - Planets - Complete for Monday
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3. Our Moon

4. Model - Revolution, Rotation and Orbit

Physics 112

Friday, October 25/19

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1. SA - U1 S3: Mathematical Analysis -> Return Monday
 2. Unit 2 - Dynamics
 3. Section 1 - Types of Forces and FBDs
 4. Force/Net Force
 5. Types of Forces
 6. Five Specific Forces
 7. [Worksheet: U2-S1 -> Weight Problems](#)
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Physics 122

Friday, October 25/19

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1. FA - Relative Velocity (RV3.1)
FA - Relative Velocity (RV3.2)
FA - Relative Velocity (RV3.3) } Due - October 25/19*
 2. Worksheet: Collisions - Elastic and Inelastic
 3. FA - 1D Explosion (CE4.2)
FA - 1D Collision (CE4.1) & Type of 1D Collision (CE4.3)
 4. 2D Collisions and Explosions
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Science 10

Friday, October 25/19

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1. FAs - Translations and Predictions -> Complete by end of class.
2. Naming Acids
3. Naming Bases
4. [Worksheet - Acids and Bases](#)

5. Neutralization Reactions
6. Reaction Rates

Naming Acids

At this level, you can recognize an acid by the fact that its formula starts with H. For example, the formula for hydrochloric acid is HCl. An exception to the rule is the formula for acetic acid which is usually given as CH_3COOH . Note that although the formula for water starts with H, it is not an acid.

Binary Acids (Anions Do Not Contain Oxygen)

All binary acids are names the same way:

1. the prefix "hydro" is used.
2. the root of the anion is used.
3. the suffix "ic" is used.
4. the word "acid" is used as the second word in the name.

HF - hydrofluoric acid
HCl - hydrochloric acid
HBr - hydrobromic acid
HI - hydroiodic acid

Oxyacids (Anions Do Contain Oxygen)

The names of these acids depend on the polyatomic ion involved. If the polyatomic ion ends in "ate", its ending will change to "ic" and "acid" will be included when naming the acid. If the polyatomic ion ends in "ite", its ending will change to "ous" and "acid" will be included when naming the acid.

ClO^- -> hypochlorite ion	HClO - hypochlorous acid
ClO_2^- -> chlorite ion	HClO_2 - chlorous acid
ClO_3^- -> chlorate ion	HClO_3 - chloric acid
ClO_4^- -> perchlorate ion	HClO_4 - perchloric acid