Science 9

Wednesday, October 30/19

http://mvhs.nbed.nb.ca/

School Pictures - Order Deadline: Thursd., Oct. 31

1. Lunch Today - SAs
2. Our Moon - Check Version of Handout
3. Model - Revolution, Rotation and Orbit - Continue

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4. More About Our Sun
5. Handouts - Guided Reading
   - Understanding the Main Ideas
6. The Motion of Our Sun
7. Comets, Asteroids and Meteors - Video Clip and Notes

Physics 112

Wednesday, October 30/19

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1. SA - U1 S3 -> Redo Thursday at Noon

2. Return:
   FA - Weight

3. Check:
   Worksheet - Free Body Diagrams

4. Review: Unit 2- Section 1: Forces and FBDs

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5. U2-S2: Newton's Laws

6. Inertia

7. Newton's First Law of Motion (Law of Inertia)

8. First Law Problems

Physics 122

Wednesday, October 30/19

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1. FA - 1D Explosion (CE4.2)
   FA - 1D Collision (CE4.1) & Type of 1D Collision (CE4.3)

2. FA - 2D Collision (CE4.4)
   FA - 2D Explosion (Ce4.5)

3. SA - Relative Velocity and 1D/2D Collisions and Explosions
   - Thursday, October 31/19
   - Six Problems: Relative Velocity (Parallel)
     Relative Velocity (Perpendicular - Boat/Plane)
     Relative Velocity (Perpendicular - Intersection)
     1D Collision & Type
     2D Collision
     2D Explosion

Science 10

Wednesday, October 30/19

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1. Return:
   FAs - Translations and Predictions

2. Questions?
   Worksheet - Neutralization Reactions

3. Reaction Rates

4. SA - Chemistry #3 - Topics
   - Date: Tuesday, Nov. 5/19

5. Review -> SA - Chemistry #3

Topics - SA: Chem #3

- be able to identify ionic compounds and molecular compounds
  - ionic compounds begin with a metallic ion or ammonium
  - molecular compounds begin with a nonmetal or metalloid
- be able to write the formulas and names for:
  (a) simple binary ionic compounds
  ie/ NaCl - sodium chloride
  (b) ionic compounds containing polyatomic ions
  ie/ $Mg(ClO_3)_2$ - magnesium chlorate
  (c) ionic compounds containing multivalent metals
  ie/ $FeBr_3$ - iron (III) bromide
  (d) ionic compounds containing multivalent metals and polyatomic ions
  ie/ Cu_3PO_4 - copper (I) phosphate
  (e) binary molecular compounds (prefixes are required for these compounds)
  ie/ P_2O_5 - diphosphorous pentoxide
  (f) binary acids (anions do not contain oxygen)
  ie/ HF - **hydro**fluor**ic acid**
  (g) oxyacids (anions do contain oxygen)
  ie/ H_2SO_4 - sulfuric acid [sulf**ate** -> sulfur**ic**]
  ie/ $HClO_2$ - chlorous acid [chlor**ite** -> chlor**ous**]
- recognize the 7 elements that form diatomic molecules (H_2, N_2, O_2, F_2, Cl_2, Br_2 and I_2), S_8, and P_4
- identify acids, bases and salts
- identify reactants and products
- be able to identify **six** types of reactions (formation, decomposition, single replacement reactions, double replacement reactions, combustion reactions, and neutralization reactions)
- be able to balance chemical reactions using numerical coefficients
- be able to translate sentences/word equations
- be able to predict products
- be able to define reaction rate
- be able to state four factors that affect reaction rates and describe how they affect reaction rates