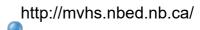
Science 9 Thursday, October 31/19



School Pictures - Order Deadline: Thursd., Oct. 31

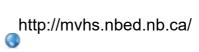
- 1. Items Owing/Updated Marks
- 2. Model Revolution, Rotation and Orbit Complete
- 3. More About Our Sun
- 4. Handouts Guided Reading- Understanding the Main Ideas
- 5. The Motion of Our Sun
- 6. Comets, Asteroids and Meteors Video Clip and Notes

Physics 112 Thursday, October 31/19

http://mvhs.nbed.nb.ca/

- 1. SA U1 S3 -> Redo at Noon
- 2. Return: FA Weight
- 3. Review: Unit 2- Section 1: Forces and FBDs
- 4. U2-S2: Newton's Laws
- 5. Inertia
- 6. Newton's First Law of Motion (Law of Inertia)
- 7. First Law Problems To Be Continued

Physics 122 Thursday, October 31/19



1. SA - Relative Velocity and 1D/2D Collisions and Explosions

Science 10 Thursday, October 31/19

http://mvhs.nbed.nb.ca/ http://mvhs-sherrard.weebly.com/

1. Return:

FAs - Translations and Predictions

- 2. SA Chemistry #3 Topics (See Next Page)- Date: <u>Tuesday, Nov. 5/19</u>
- 3. Review -> SA Chemistry #3 (*Reaction Rates will be on SA)
- 4. Extra Practice Mixed Nomenclature Practice
 Practice Worksheet Naming Acids
 Word Equations
 Chemical Reactions and Chemical Equations
 More Predicting Products of Chemical Reactions

Topics - SA: Chem #3

- be able to identify ionic compounds and molecular compounds
 - ionic compounds begin with a metallic ion or ammonium
 - molecular compounds begin with a nonmetal or metalloid
- be able to write the formulas and names for:
 - (a) simple binary ionic compounds ie/ NaCl sodium chloride
 - (b) ionic compounds containing polyatomic ions ie/ Mg(ClO₃)₂ magnesium chlorate
 - (c) ionic compounds containing multivalent metals ie/ FeBr₃ iron (III) bromide
 - (d) ionic compounds containing multivalent metals and polyatomic ions ie/ Cu₃PO₄ copper (I) phosphate
 - (e) binary molecular compounds (prefixes are required for these compounds)
 - ie/ P₂O₅ diphosphorous pentoxide
 - (f) binary acids (anions do not contain oxygen) ie/ HF hydrofluoric acid
 - (g) oxyacids (anions do contain oxygen)

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ie/ H<sub>2</sub>SO<sub>4</sub> - sulfuric acid [sulfate -> sulfuric]
ie/ HClO<sub>2</sub> - chlorous acid [chlorite -> chlorous]
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- recognize the 7 elements that form diatomic molecules (H_2 , N_2 , O_2 , F_2 , Cl_2 , Br_2 and I_2), S_8 , and P_4
- identify acids, bases and salts
- identify reactants and products
- be able to identify **six** types of reactions (formation, decomposition, single replacement reactions, double replacement reactions, combustion reactions, and neutralization reactions)
- be able to balance chemical reactions using numerical coefficients
- be able to translate sentences/word equations
- be able to predict products
- be able to define reaction rate
- be able to state four factors that affect reaction rates and describe how they affect reaction rates