

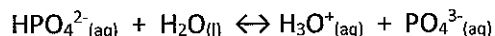
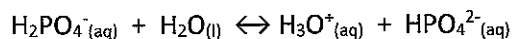
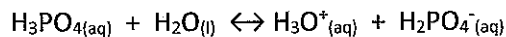
Monoprotic – an acid that only donates one H^+ ion

Diprotic – donates two H^+ ions

Triprotic – donates three H^+ ions

Polyprotic – any acid that has more than one ionizable H^+ . They ionize in steps.

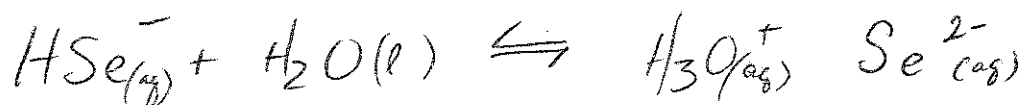
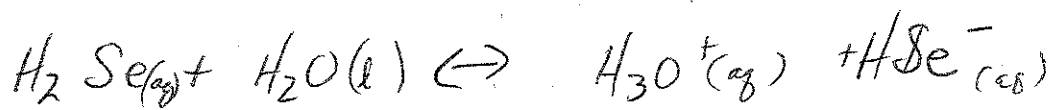
For example, $H_3PO_{4(aq)}$.



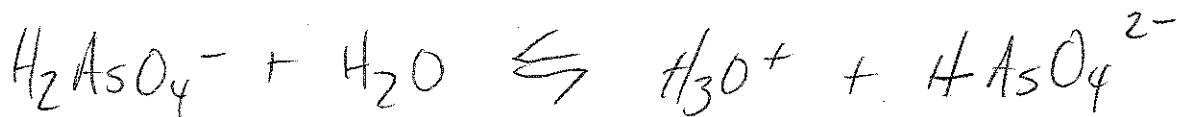
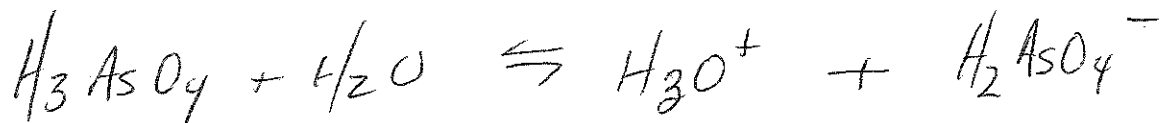
The ability for a solution to donate H^+ ions depends on its polarity. If a molecule is nonpolar, the H^+ ions will not ionize.

3. Write the steps in the complete ionization of the following polyprotic acids.

a. H_2Se



b. H_3AsO_4



c. H_2SO_3

